

Common Polyatomic Ions							
+1		-1		-2		-3	
$\text{NH}_4^+$	ammonium	$\text{C}_2\text{H}_3\text{O}_2^-$	acetate	$\text{CO}_3^{2-}$	carbonate	$\text{PO}_4^{3-}$	phosphate
$\text{H}_3\text{O}^+$	hydronium	$\text{ClO}^-$	hypochlorite	$\text{CrO}_4^{2-}$	chromate	$\text{PO}_3^{3-}$	phosphite
		$\text{ClO}_2^-$	chlorite	$\text{Cr}_2\text{O}_7^{2-}$	dichromate		
		$\text{ClO}_3^-$	chlorate	$\text{SO}_4^{2-}$	sulfate		
		$\text{ClO}_4^-$	perchlorate	$\text{SO}_3^{2-}$	sulfite		
		$\text{CN}^-$	cyanide	$\text{O}_2^{2-}$	peroxide		
		$\text{NO}_3^-$	nitrate	$\text{C}_2\text{O}_4^{2-}$	oxalate		
		$\text{NO}_2^-$	nitrite				
		$\text{HCO}_3^-$	hydrogen carbonate (bicarbonate)				
		$\text{OH}^-$	hydroxide				
		$\text{MnO}_4^-$	permanganate				

Common Metal Ions			Prefixes Used in Naming Binary Molecular Compounds	
Ion	Systematic Name	Common Name	Prefix	Number
$\text{Fe}^{2+}$	iron (II)	ferrous	mono	1
$\text{Fe}^{3+}$	iron (III)	ferric	di	2
$\text{Cu}^+$	copper (I)	cuprous	tri	3
$\text{Cu}^{2+}$	copper (II)	cupric	tetra	4
$\text{Pb}^{2+}$	lead (II)	plumbous	penta	5
$\text{Pb}^{4+}$	lead (IV)	plumbic	hexa	6
$\text{Cr}^{2+}$	chromium (II)	chromous	hepta	7
$\text{Cr}^{3+}$	chromium (III)	chromic	octa	8
$\text{Sn}^{2+}$	tin (II)	stannous	nona	9
$\text{Sn}^{4+}$	tin (IV)	stannic	deca	10
$\text{Co}^{2+}$	cobalt (II)	cobaltous		
$\text{Co}^{3+}$	cobalt (III)	cobaltic		
$\text{Hg}_2^{2+}$	mercury (I)	mercurous		
$\text{Hg}^{2+}$	mercury (II)	mercuric		

\*\* ALWAYS:  $\text{Zn}^{2+}$ ,  $\text{Ag}^+$ ,  $\text{Cd}^{2+}$  \*\*